



# IJRASET

International Journal For Research in  
Applied Science and Engineering Technology



---

# INTERNATIONAL JOURNAL FOR RESEARCH

IN APPLIED SCIENCE & ENGINEERING TECHNOLOGY

---

**Volume: 5      Issue: XII      Month of publication: December 2017**

**DOI:**

[www.ijraset.com](http://www.ijraset.com)

Call:  08813907089

E-mail ID: [ijraset@gmail.com](mailto:ijraset@gmail.com)

# Iridium (III) Catalyzed Oxidation of D-Sorbitol By Sodium N-Chloro-P-Toluenesulfonamide in Perchloric Acid Medium: Mechanistic Approach and Kinetic Modelling

Abhishek Verma<sup>1</sup>, Jyoti Pandey<sup>2</sup>, Reena Patel<sup>3</sup>, Sheila Srivastava<sup>4</sup>

<sup>1, 2, 3</sup> Department of Applied Chemistry, BBAU, Lucknow-226025, U.P., India

<sup>4</sup> Chemical laboratories, Feroze Gandhi College, Raebareli-229001, U.P., India

**Abstract:** The present paper deals with the kinetics and mechanism of homogeneously Ir(III) chloride catalyzed oxidation of D-sorbitol by chloramine-T [CAT] in perchloric acid medium in the temperature range of 30 to 45 °C. The reaction is carried out in the presence of mercuric acetate as a scavenger for chloride ion. The experimental results show first order kinetics with respect to the oxidant [CAT] and catalyst [Ir(III)] while zero order with respect to substrate, i.e., D-Sorbitol was observed. The reaction shows negligible effect of [Hg(OAc)<sub>2</sub>], [H<sup>+</sup>] and ionic strength of the medium. Chloride ion positively influences the rate of reaction. A transient complex, formed between Ir(III) and chloramine-T in a slow and rate determining step, reacts with the D-sorbitol to give the product in a series of fast steps, which lead to regeneration of the catalyst. The reaction between chloramine-T and D-Sorbitol in acid medium shows 2:1 stoichiometry. To calculate activation parameters, the reactions have been studied at four different temperatures between 30 to 45 °C. A mechanism involving the complex formation between catalyst and oxidant has been proposed. D-Gluconic acid has been identified chromatographically and spectroscopically as the final product of oxidation of D-Sorbitol. Based on the kinetic data, reaction stoichiometry and product analysis, a reaction mechanism has been proposed and rate law has been derived.

**Key Words:** Kinetics, mechanism, Ir(III) catalysis, D-sorbitol, chloramine-T, acidic medium.

## I. INTRODUCTION

The sodium salts of arylhalosulfonamides generally known as organic haloamines have attracted the attention of chemists, as versatile redox titrants [1]. The miscellaneous chemical behaviour of organic haloamines is accredited in general to their ability to act as halonium cations, hypohalites, N-anions which act both as bases and nucleophiles and nitrenoids in limiting cases [2]. As a result, these compounds react with a wide range of functional groups to cause numerous kinds of molecular changes. Aromatic N-haloamines act as mild oxidants in both acid and alkaline solutions due to the presence of strongly polarized N-linked halogen in +1 state [3,4]. A prominent member of this group Sodium N-chloro-p-toluenesulfonamide or Chloramine-T (CAT; p-CH<sub>3</sub>C<sub>6</sub>H<sub>4</sub>SO<sub>2</sub>NCINa.3H<sub>2</sub>O) is used as halogenating and oxidizing agent [4-6]. CAT is a byproduct in the manufacture of saccharin [7]. Depending upon the pH of the medium, it forms various oxidizing species and thus shows a variety of kinetic results [8,9]. Several researchers have studied the oxidizing behavior of CAT [10-13] and numerous studies focus on the mechanistic aspects of the redox reactions in acidic media. In most of the studies one of the species, RNHCl (R = CH<sub>3</sub>C<sub>6</sub>H<sub>4</sub>SO<sub>2</sub>), HOCl, or H<sub>2</sub>OCl<sup>+</sup>, has been considered as the reactive species [14]. It can behave as both electrophile and nucleophile depending on the reaction conditions. A polyhydric alcohol or polyalcohol is a hydrogenated form of carbohydrate, whose carbonyl group (aldehyde or ketone, reducing sugar) has been reduced to primary and secondary hydroxyl group. D-Sorbitol, a polyhydric alcohol with about half the sweetness of sucrose. D-Sorbitol occurs naturally and is also produced synthetically from glucose. It was formerly used as a diuretic and may still be used as a laxative and in irrigating solutions for some surgical procedures. It is also used in many manufacturing processes, as a pharmaceutical aid, and in several research applications. Commonly the study of oxidation of polyhydric alcohols has great importance in sugar chemistry. Various investigations on oxidation involving polyhydric alcohols and various oxidants i.e., cerium(VI) [15-17], potassium bromate [18], N-Bromosuccinimide [19] in the presence/absence of catalyst in acidic medium have been reported by various workers. However, there are only few reports available in the literature for the oxidation of D-Sorbitol [20].



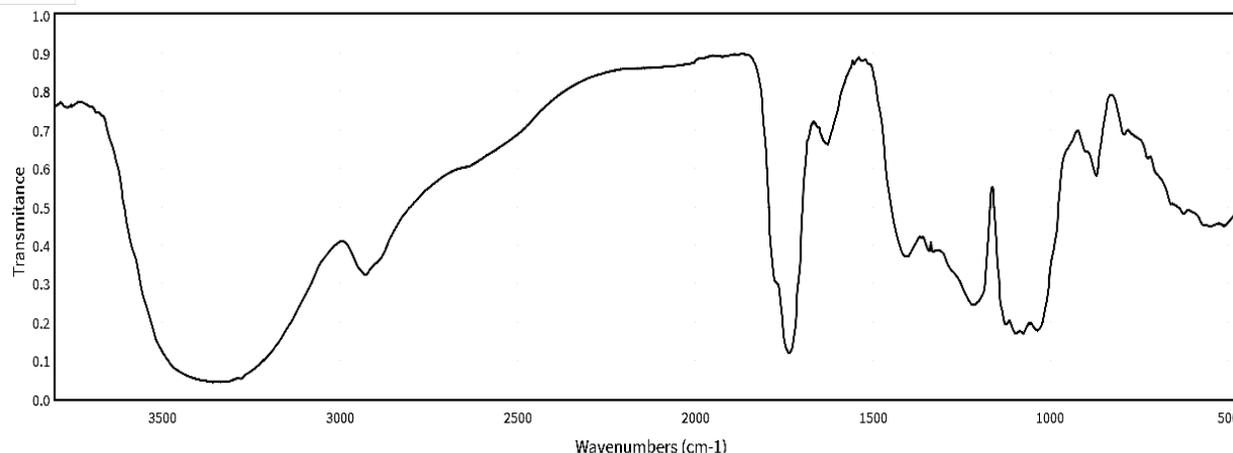


Fig. 1. IR-Spectra of the main product D-gluconic acid

The C-O stretching vibration observed at 1237 cm<sup>-1</sup> and C=O stretching at 1725 cm<sup>-1</sup>. The C-O-H in plane bending appears at 1165cm<sup>-1</sup> confirms the presence of – OH group to the carbon frame .The C-C stretching vibration appears at 1020cm<sup>-1</sup>.The C-H stretching vibration of ν (C-H) appears at 2943 cm<sup>-1</sup>.

### III. RESULT AND DISCUSSION

Kinetics of Ir(III) catalyzed oxidation of D-Sorbitol by chloramine-T in acidic medium was investigated at 35 °C. The kinetic results were collected at several initial concentrations (Table1). The order of reaction with respect to each reactant was determined by varying the concentration of oxidant, D-Sorbitol, Ir(III) chloride (Table 1), H<sup>+</sup> ions, [Cl<sup>-</sup>] and mercuric acetate one by one in different sets keeping concentration of all other reactants constant at constant temperature 35 °C. In each kinetic runs, the initial rate (i.e., -dc/dt) of the reaction was determined from the slope of the tangent drawn at a fixed concentration of chloramine-T except for the chloramine-T variation in which the slope of the tangent was drawn at fixed time. The first order reaction rate constant (k<sub>1</sub>) for the variation of all the reagents were calculated:

$$k_1 = \frac{-dc/dt}{[RNHCl]^*}$$

Where [RNHCl]\* denotes the [RNHCl] at which (-dc/dt) was determined.

Table 1. Effect of variation of oxidant, D-sorbitol, Ir(III) at 35°C

[Oxidant] x 10 <sup>3</sup> M (Chloramine-T)	[Substrate]x 10 <sup>2</sup> M (D- sorbitol)	[Ir(III)] x 10 <sup>5</sup> M	(-dc/dt)x10 <sup>7</sup> ML <sup>-1</sup> s <sup>-1</sup>
0.83	2.00	8.02	3.45
1.00	2.00	8.02	4.07
1.25	2.00	8.02	4.95
1.67	2.00	8.02	6.65
2.50	2.00	8.02	10.12
5.00	2.00	8.02	20.28
1.00	0.33	8.02	3.72
1.00	0.40	8.02	4.47
1.00	0.50	8.02	3.95
1.00	0.66	8.02	3.56
1.00	1.00	8.02	4.28
1.00	2.00	8.02	4.07

1.00	2.00	2.67	1.35
1.00	2.00	4.01	2.05
1.00	2.00	5.34	2.66
1.00	2.00	6.67	3.42
1.00	2.00	8.02	4.07
1.00	2.00	9.35	5.06

Solution conditions:  $[Hg(OAc)_2] = 1.25 \times 10^{-3} M$ ,  $[HClO_4] = 1.00 \times 10^{-3} M$ ,  $[KCl] = 1.00 \times 10^{-3} M$ .

A plot of  $(-dc/dt)$  versus  $[Chloramine-T]$  was linear with the slope value (0.97) near unity, which further confirms first order dependence of reaction on chloramine-T (Fig. 2). The first order kinetics w.r.t.  $[Chloramine-T]$  was also confirmed by 'least-square method' (Fig. 3). A plot of  $(-dc/dt)$  versus  $[Ir(III)]$  gives a slope which is close to the average value of first order rate constant at 35 °C (Fig. 4). Insignificant effect on the rate was observed on increasing the concentration of the substrate, indicating zero order in substrate *i.e.* D-Sorbitol (Table1).

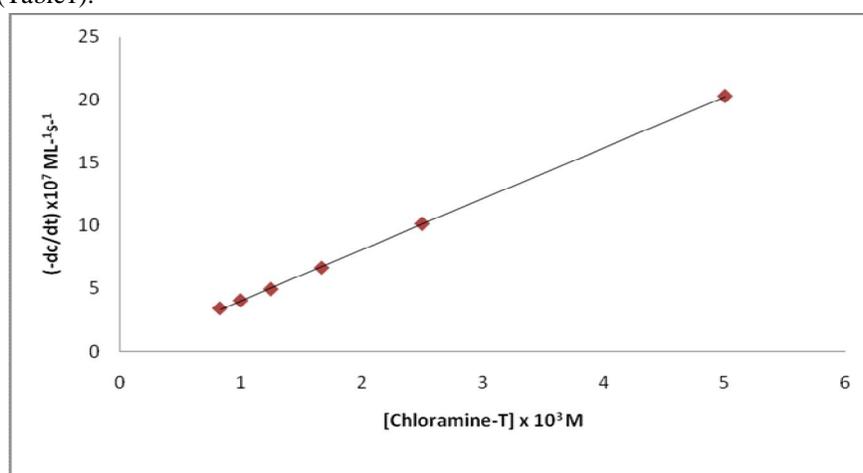


Fig. 2. Plot between  $(-dc/dt)$  and  $[Chloramine-T]$  for the oxidation of D-sorbitol at 35 °C.  $[Ir(III)] = 8.02 \times 10^{-5} M$ ,  $[D-sorbitol] = 2.00 \times 10^{-2}$ ,  $[Hg(OAc)_2] = 1.25 \times 10^{-3} M$ ,  $[HClO_4] = 1.00 \times 10^{-3} M$ ,  $[KCl] = 1.00 \times 10^{-3} M$ .

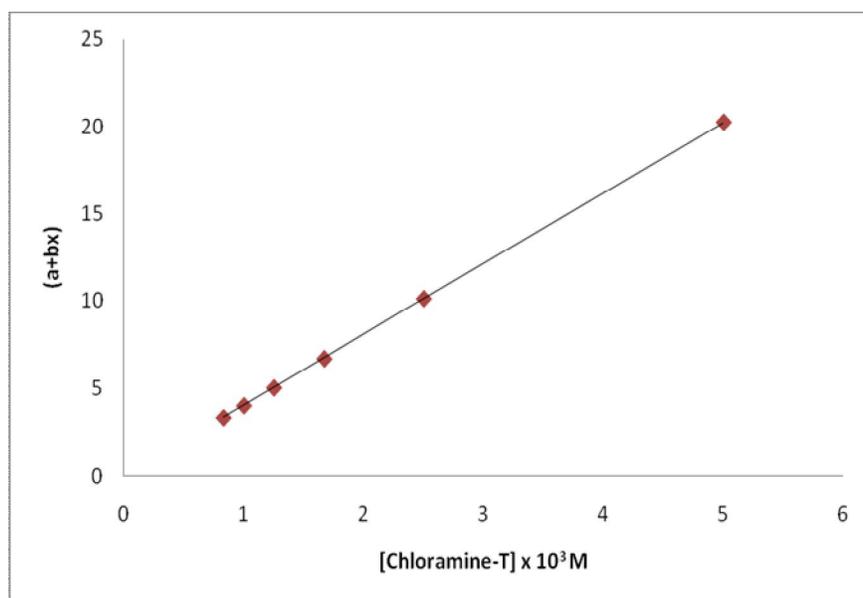


Fig. 3. Plot between  $(a + bx)$  and  $[Chloramine-T]$  for the oxidation of D-sorbitol at 35 °C.  $[Ir(III)] = 8.02 \times 10^{-5} M$ ,  $[D-sorbitol] = 2.00 \times 10^{-2}$ ,  $[Hg(OAc)_2] = 1.25 \times 10^{-3} M$ ,  $[HClO_4] = 1.00 \times 10^{-3} M$ ,  $[KCl] = 1.00 \times 10^{-3} M$ .

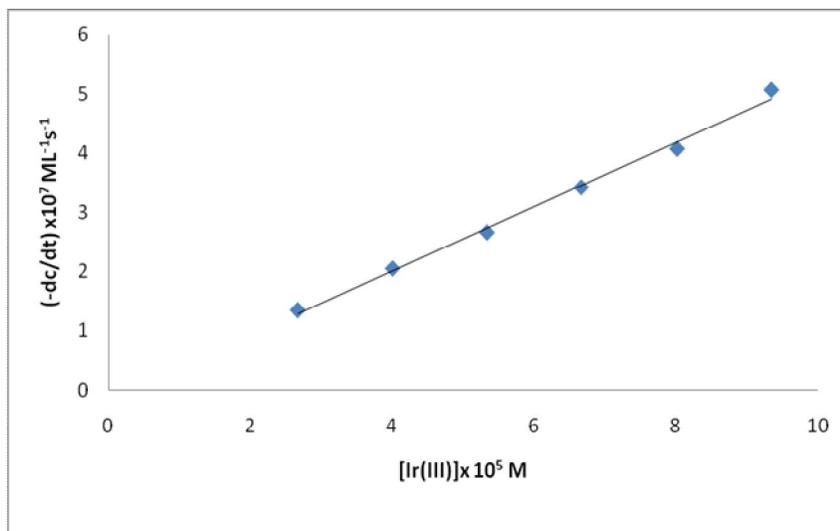


Fig. 4. Plot between  $(-dc/dt)$  and  $[Ir(III)]$  for the oxidation of D-sorbitol at 35°C.  $[D\text{-sorbitol}] = 2.00 \times 10^{-2} \text{ M}$ ,  $[\text{Chloramine-T}] = 1.00 \times 10^{-3}$ ,  $[\text{Hg}(\text{OAc})_2] = 1.25 \times 10^{-3} \text{ M}$ ,  $[\text{HClO}_4] = 1.00 \times 10^{-3} \text{ M}$ ,  $[\text{KCl}] = 1.00 \times 10^{-3} \text{ M}$ .

Variation of  $[\text{KCl}]$  concentration shows positive effect on reaction rate (Fig. 5) (Table 2). Negligible effect of mercuric acetate eliminate the probability of its involvement either as a catalyst or as an oxidant because it does not help the reaction proceed without  $[\text{CAT}]$ . Hence, the function of mercuric acetate is to act as scavenger for any chloride ion formed in the reaction [28]. It helps to eliminate the parallel oxidation by  $\text{Cl}_2$  which would have been formed as a result of interaction between  $\text{Cl}^-$  and  $\text{RNHCl}$  ion. Experimental data indicate negligible effect of ionic strength of the medium on the rate (affected by addition of  $\text{NaClO}_4$ ). In acidic solution of chloramine-T quick formation of  $\text{RNHCl}$  has been reported [29]. The reaction is unaffected by  $\text{H}^+$  concentration (Table 2)

Table 2. Effect of variation of  $\text{HClO}_4$ ,  $\text{KCl}$  and  $\text{NaClO}_4$  at 35°C

$[\text{HClO}_4] \times 10^3 \text{ M}$	$[\text{KCl}] \times 10^3 \text{ M}$	$\text{NaClO}_4 \times 10^3 \text{ M}$	$(-dc/dt) \times 10^7 \text{ ML}^{-1} \text{ s}^{-1}$
0.83	1.00	1.00	3.42
1.00	1.00	1.00	4.07
1.25	1.00	1.00	2.96
1.67	1.00	1.00	3.98
2.50	1.00	1.00	3.72
5.00	1.00	1.00	3.50
1.00	0.83	1.00	3.64
1.00	1.00	1.00	4.07
1.00	1.25	1.00	4.45
1.00	1.67	1.00	4.92
1.00	2.50	1.00	5.35
1.00	5.00	1.00	5.96
1.00	1.00	0.83	4.24
1.00	1.00	1.00	4.07
1.00	1.00	1.25	3.56
1.00	1.00	1.67	3.25
1.00	1.00	2.50	3.92
1.00	1.00	5.00	3.36

Solution conditions:  $[\text{Ir}(\text{III})] = 8.02 \times 10^{-5} \text{ M}$ ,  $[\text{Chloramine-T}] = 1.00 \times 10^{-3} \text{ M}$ ,  $[\text{D-sorbitol}] = 2.00 \times 10^{-2}$ ,  $[\text{Hg}(\text{OAc})_2] = 1.25 \times 10^{-3} \text{ M}$ .

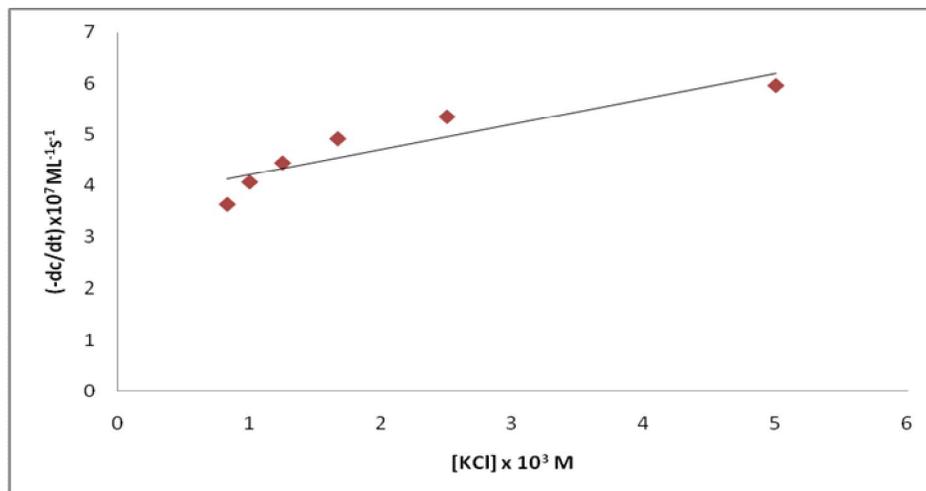


Fig. 5. Plot between  $(-dc/dt)$  and  $[KCl]$  for the oxidation of D- sorbitol at 35 °C.  $[Ir(III)] = 8.02 \times 10^{-5} M$ ,  $[D\text{-sorbitol}] = 2.00 \times 10^{-2} M$ ,  $[Chloramine\text{-}T] = 1.00 \times 10^{-3}$ ,  $[Hg(OAc)_2] = 1.25 \times 10^{-3} M$ ,  $[HClO_4] = 1.00 \times 10^{-3} M$ .

The kinetic measurements were also taken in the temperature range of 30-45 °C and the specific rate constants were obtained at 30 °C, 35 °C, 40 °C and 45 °C. These specific rate constants were used to draw a plot of  $\log k$  versus  $1/T$  (Fig. 6) which was linear. The various activation parameters were calculated from the slope of curves obtained from the Fig. 6. The value of energy of Activation ( $E_a$ ), Arrhenius factor ( $A$ ), entropy of activation ( $\Delta S^*$ ), free energy of activation ( $\Delta G^*$ ) and enthalpy of activation ( $\Delta H^*$ ) were calculated from rate measurement and these values have been provided in Table-3. Moderate  $\Delta H^*$  and  $\Delta S^*$  values are favorable for electron transfer reaction. The value of  $\Delta H^*$  was due to energy of solution changes in transition state. The high positive value of  $\Delta G^*$  represents highly solvated transition state. The negative value of  $\Delta S^*$  indicates that the intermediate complex is more ordered than the reactants so the formation of activated complex occurs with reduction in the degree of freedom [30]. The observed modest enthalpy of activation and higher rate constant for the slow step shows that oxidation presumably occurs by means of an inner sphere mechanism [31].

Table 3. Activation parameters for the oxidation of D-sorbitol

Parameters	Temperature(°C)	D-sorbitol
$K_1 \times 10^4 s^{-1}$	30	3.15
$K_1 \times 10^4 s^{-1}$	35	4.07
$K_1 \times 10^4 s^{-1}$	40	6.27
$K_1 \times 10^4 s^{-1}$	45	8.12
Log A	--	7.88
$E_a$ (kJ mol <sup>-1</sup> )	35	42.90
$\Delta G^*$ (kJ mol <sup>-1</sup> )	35	70.51
$\Delta H^*$ (kJ mol <sup>-1</sup> )	35	40.33
$\Delta S^*$ (JK <sup>-1</sup> mol <sup>-1</sup> )	35	-98.05

Solution conditions:  $[Ir(III)] = 8.02 \times 10^{-5} M$ ,  $[Chloramine\text{-}T] = 1.00 \times 10^{-3} M$ ,  $[D\text{-Sorbitol}] = 2.00 \times 10^{-2}$ ,  $[Hg(OAc)_2] = 1.25 \times 10^{-3} M$ ,  $[HClO_4] = 1.00 \times 10^{-3} M$ ,  $[KCl] = 1.00 \times 10^{-3} M$ .

This conclusion is supported by earlier observations. The activation parameters evaluated for the catalyzed reaction explain the catalytic effect on the reaction. Kinetic observations show that the reaction under investigation is complex reaction, which usually takes place in more than one step.

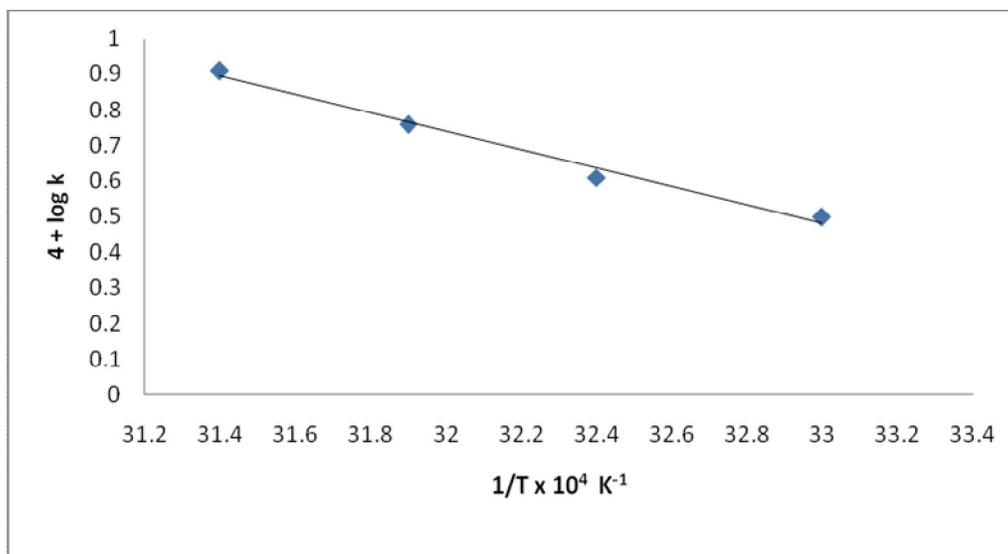


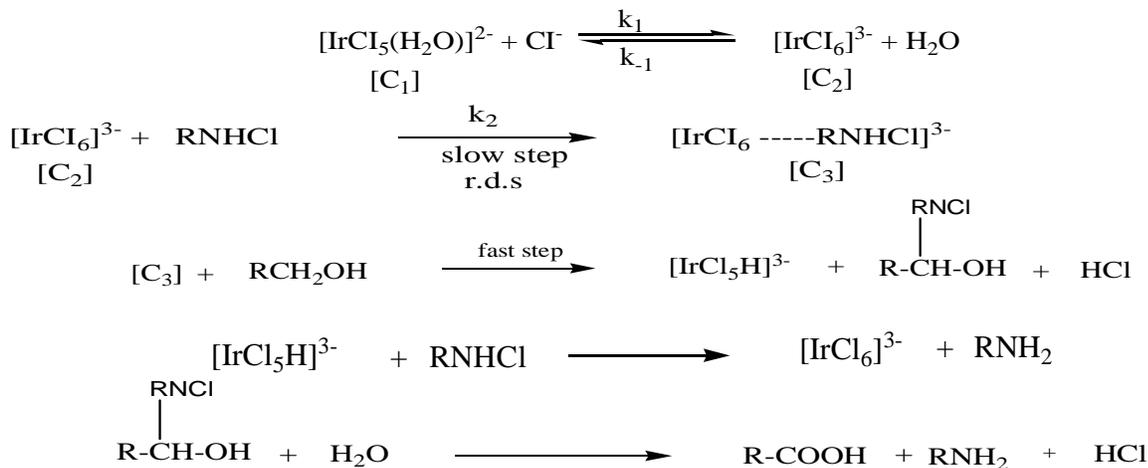
Fig. 6. Arrhenius plot for the oxidation of D-sorbitol.  $[\text{Ir (III)}] = 8.02 \times 10^{-5} \text{ M}$ ,  $[\text{Chloramine-T}] = 1.00 \times 10^{-3} \text{ M}$ ,  $[\text{D-sorbitol}] = 2.00 \times 10^{-2}$ ,  $[\text{Hg(OAc)}_2] = 1.25 \times 10^{-3} \text{ M}$ ,  $[\text{HClO}_4] = 1.00 \times 10^{-3} \text{ M}$ ,  $[\text{KCl}] = 1.00 \times 10^{-3} \text{ M}$ .

#### A. Mechanism and derivation of rate law

The acidic solution of Iridium chloride exists as  $[\text{IrCl}_6]^{3-}$ . It has also been reported that  $[\text{IrCl}_6]^{3-}$  is involved in equilibrium as follows [32] :



Thus either  $[\text{IrCl}_6]^{3-}$  or  $[\text{IrCl}_5(\text{H}_2\text{O})]^{2-}$  may act as catalytic species [33]. If  $[\text{IrCl}_5(\text{H}_2\text{O})]^{2-}$  is taken as catalytic species the rate law would require negative effect of chloride ion contrary to the positive effect of chloride ion on the oxidation rate observed by us. Hence the only choice is  $[\text{IrCl}_6]^{3-}$  which when assumed as reactive species of Iridium trichloride in acidic medium, explains the positive effect of chloride ion. The kinetic results reported in table 1, 2, 3 along with the above discussion lead us to suggest the following reaction scheme:



Where,

$\text{R} = \text{CH}_2\text{OH}(\text{CHOH})_n$ , for D-Sorbitol ( $n = 4$ )

[P] = Product i.e. D-Gluconic acid

Now, Considering the above reaction steps and applying the steady-state treatment with reasonable approximation, the rate law may be written as

$$\text{Rate} = \frac{-d[\text{RNHCl}]}{dt} = k_2 [\text{C}_2] [\text{RNHCl}] \quad (i)$$

$$[\text{Ir (III)}]_T = [\text{C}_1] + [\text{C}_2] \quad (ii)$$

$$\frac{d[C_1]}{dt} = k_{-1} [C_2] - k_1 [C_1][Cl^-] \quad (iii)$$

$$[C_1] = \frac{k_{-1} [C_2]}{k_1 [Cl^-]} \quad (iv)$$

$$[C_1] = \frac{[C_2]}{K_1 [Cl^-]} \quad (v)$$

(where  $K_1 = k_1/k_{-1}$ )

Putting the value of  $[C_1]$  in equation (ii), we get :

$$\begin{aligned} [Ir(III)]_T &= [C_1] + [C_2] \\ [Ir(III)]_T &= \frac{[C_2]}{K_1 [Cl^-]} + [C_2] \\ &= \left[ \frac{[C_2] + K_1 [Cl^-][C_2]}{K_1 [Cl^-]} \right] \\ &= [C_2] \left[ \frac{1 + K_1 [Cl^-]}{K_1 [Cl^-]} \right] \end{aligned}$$

This gives , 
$$[C_2] = \frac{[Ir(III)]_T K_1 [Cl^-]}{1 + K_1 [Cl^-]}$$

Putting the value of  $C_2$  in equation (i), we get,

$$Rate = \frac{K_1 k_2 [Ir(III)]_T [Cl^-] [RNHCl]}{1 + K_1 [Cl^-]}$$

The above equation is final rate law which well explains the experimental first-order kinetics with respect to [chloramine-T] and [Ir(III)] on the rate of oxidation and also order with respect to  $[Cl^-]$  is fractional positive.

#### IV. COMPARATIVE STUDIES

The results of present study of Ir(III) catalyzed oxidation of D-sorbitol by chloramine-T in acidic medium can be compared with the results for the Ru(III) catalyzed oxidation of polyhydric alcohols (D-mannitol and D-sorbitol) by  $KBrO_3$  in alkaline medium [20]. The present study being similar in order with respect to oxidant and catalyst but shows significant change as far as order with respect to  $Cl^-$  ion is concerned. The observed positive effect of  $[Cl^-]$  on the rate of oxidation in present study is contrary to the reported negative effect of  $[Cl^-]$  in the oxidation of polyhydric alcohols ( D-mannitol and D-sorbitol) using Ru(III) chloride as a catalyst [20]. The most reactive species of Ir(III) chloride is  $[IrCl_6]^{3-}$  in acidic medium in present study while  $[RuCl_2(H_2O)_3(OH)]$  reported as most reactive species of Ru(III) chloride in alkaline medium [20]. The stoichiometric analysis of present study shows 2 mole of oxidant[RNHCl] consumes 1 mole of substrate i.e. D-sorbitol, On the other hand in Ru(III) catalyzed reaction one mole of oxidant  $[KBrO_3]$  consumes one mole of substrate i.e. polyhydric alcohols [20]. On the basis of facts mentioned above, it can be inferred that the present study differs in some respects from the study reported for Ru(III) chloride catalyzed oxidation.

#### V. CONCLUSION

The following conclusions can be derived in the present study of Ir(III) catalyzed oxidation of D-sorbitol by chloramine-T in acidic medium. (a) Among the various species of Ir(III) in acidic medium,  $[IrCl_6]^{3-}$  is considered as the reactive species while (b) RNHCl is the reactive species of Chloramine-T in acidic medium. (c) In the absence of catalyst oxidation of D-sorbitol by Chloramine-T is very sluggish, but it becomes facile in the presence of Ir(III) catalyst. (d) The stoichiometry of the reaction was found to be 2:1 and

the oxidation products of were identified (e) Activation parameters were computed from the Arrhenius plot. (f) The observed results have been explained by a plausible mechanism and the related rate law has been deduced. It can be concluded that Ir(III) chloride act as an efficient catalyst for the oxidation of D-sorbitol by Chloramine-T in acidic medium.

## VI. ACKNOWLEDGEMENT

The first author thankfully acknowledges the CSIR, New Delhi, India for providing financial assistance in the form of Research Fellowship.

## REFERENCES

- [1] Puttaswamy, J.P. Shubha, R.V. and Jagadeesh, "Ruthenium(III)-catalyzed oxidative cleavage of p-aminoazobenzene by chloramine-B in alkaline medium and uncatalyzed reaction in acid medium: Spectrophotometric kinetic and mechanistic study", *Tans. Met. Chem.*, 32, 991 (2007).
- [2] Jennings, V.J., "Analytical Applications of Chloramine-T", *Critical Rev. Anal-Chem.*, 3(4), 407(1974).
- [3] Kolvari, A. Choghamarani, A.G. Salehi, P. Shirini, F. and Zolfigol, M A., "Applications of N- halo reagents in organic synthesis", *J. Iran. Chem. Soc.*, 4(2), 126 (2007).
- [4] Umemoto, T. Fukami, S. Tomizava, G. Harasava, K. Kavada, K. and Tomita, K., "Formal electrophilic reactions of fluorine presumably involve alternative Mechanisms", *J. Am. Chem. Soc.*, 112, 8563 (1990).
- [5] Ramananda, M. G. and Ananda, S., "Kinetics and mechanism of oxidation of o-toluidine by sodium -N-chloro benzene sulphonamide in acid medium", *Asian J. Chem.*, 11(4), 1348 (1999).
- [6] Singh, A. K. Negi, R. Jain, B. Katre, Y. Singh, S.P. and Sharma V. K., "Kinetics and Mechanism of Ru(III)-Catalyzed Oxidation of Paracetamol by Chloramine-T in Aqueous Acidic Medium", *Catal. Lett.*, 132, 285 (2009).
- [7] Campbell, M.M. and Johnson, G., "Chloramine-T and related N-halogeno-N-metallo reagents" *Chem. Rev.*, 8, 65 (1978).
- [8] Singh, A.K. Negi, R. Katre, Y. and Singh, S.P. "Mechanistic study of novel oxidation of paracetamol by chloramine-T using micro-amount of chloro-complex of Ir(III) as a homogeneous catalyst in acidic medium", *J. Mol. Catal A: Chem.*, 302, 36 (2009).
- [9] Puttaswamy, Jagadeesh, R.V. and Vaz, N., "Ru(III)-catalyzed oxidation of some N-heterocycles by chloramine-T in hydrochloric acid medium: a kinetic and mechanistic study", *J. Mol. Catal A*, 229, 211 (2009).
- [10] Filler, R., "Oxidations and Dehydrogenations with N-Bromosuccinimide and related N- Haloimides", *Chem. Rev.*, 63, 21 (1963)
- [11] Wadhvani, M. and Jain, S., "Kinetic study of oxidation of some monosaccharides by Chloramine-T induced by visible light source", *Res. J. Chem. Sci.*, 4(7), 61 (2014).
- [12] Veena, M.S. Prashanth, M.K. Kumar, K.Y. Muarlidhara, H.B. and Nayaka, Y.A., "Kinetics and mechanistic study of oxidation of amoxicillin by chloramine-T in acid medium", *J. Chil. Chem. Soc.*, 60(3), 3063 (2015)
- [13] Wadhvani, M. and Jain, S., "Ultraviolet Light Induced Oxidation of Malonic Acid by Chloramine-T in an Aqueous Acidic Medium: A Kinetic Study", *Ind. J. Adv. Chem. Sci.*, 40(1), 36 (2016).
- [14] Singh, A.K. Negi, R. Jain, B. Katre, Y. Singh, S.P. and Sharma V.K., "Pd (II) catalyzed oxidative degradation of paracetamol by chloramine-T in acidic and alkaline media", *Ind. Eng. Chem. Res.*, 50, 8407 (2011).
- [15] Rajeswari, V. Hosahalli, Anita P. SavanurSharanappat, Nandibeworr and Shivamrti A., "Ru(III) mediated oxidation of D-Mannitol by cerium (IV) in aqueous Sulphuric acid medium: a kinetic & mechanistic approach" *wiley inter science*, 10J, 1002 (2010)
- [16] Agarwal, A. Sharma, G. Khandelwal, C.L. and Sharma P.D., "Kinetics and Mechanism of Oxidation of Mannitol by Sulphato-cerium(IV) Species in Aqueous Acid Medium", *Inorg. React. Mech.*, 4(3), 233(2002).
- [17] Aparna, R. and Das, A.K., "Studies on kinetics and mechanism of iridium(III) catalysed cerium(IV) oxidation of D-mannitol and D-glucose in aqueous acid media", *Indian J. chem.*, 41(A), 2468 (2002).
- [18] Srivastava, S. Srivastava, P., "Kinetics and mechanism of Oxidation of D- Mannitol by Potassium bromate in aqueous acidic medium", *Pelg. res. Lib.*, 1(1), 13 (2010).
- [19] Singh, B. Singh, D. Singh, S. and Kumar, A., "Kinetics and Mechanism of Ru(III) catalysis in the oxidation of Erythritol and Dulcitol by NBS", *Trans. Met. chem.*, 16, 610-613 (1991)
- [20] Srivastava, S. and Singh, A.L., "Mechanistic studies of oxidation of D-sorbitol and D-mannitol by an alkaline solution of potassium bromate in the presence of chloro complex of Ru(III) as homogeneous catalyst", *Oxid. Communi.*, 36(1), 41(2013).
- [21] Puttaswamy, Vaz, N., Ruthenium(III) and osmium(VIII)-catalyzed oxidation of 2-thiouracil by bromamine-B in acid and alkaline media: a kinetic and mechanistic study. *Trans. Met. Chem.*, 28, 409-417(2003).
- [22] Hiremath, C.V. Kiran, T.S. and Nandibeworr, S.T., "Os(VIII)/Ru(III) catalysed oxidation of aspirin drug by a new oxidant, diperiodatoargentate(III) in aqueous alkaline medium: A comparative kinetic study", *J. Mol. Catal A. Chem.*, 248, 163 (2006).
- [23] Uma, V. Sethuram, B. and Rao, T.N., "Kinetics of oxidation of isopropanol by  $IO_4^-$  catalyzed by metal ion", *React. Kinet. Catal. Lett.*, 18, 283(1981)
- [24] Song, W. Zhao, R. Jiang, Q., "Catalyzed oxidation of Tetrahydrofurfuryl alcohol by Cerium(IV) in aqueous sulphuric acid medium.", *Acta. Phys. Chim. Sin.*, 21(08), 929(2005).
- [25] Singh, A.K. Rahmani, S. Singh B. Singh, R.K. and Singh M., "Mechanism of Ir(III)-catalysed and Hg(II)-co-catalysed oxidation of reducing sugars by N-bromoacetamide in acidic medium", *J. Phys. Org. Chem.*, 17(3), 249 (2004).
- [26] Bai, W. Zhai, Y. Liu, H. Lu, H. Song, W., "Kinetics and mechanism of iridium(III) catalyzed oxidation of ethylene glycol by cerium(IV) in sulfuric acid media" *Ind. J. Chem.*, 47A, 1194 (2008).
- [27] Singh, A.K. Sachdev, N. Srivastava, A. Jain, B. and Khtre, Y., "Oxidation of d-glucose by N-bromophthalimide in the presence of chlorocomplex of iridium (III): a kinetic and mechanistic study" *Res. Chem. Intermed.*, 38, 507 (2012).
- [28] Venkatasubramanian, N. and Thiagarajan, V., "The mechanism of oxidation of alcohols by bromine" *Tetra. Lett.* 9, 1711 (1968).



- [29] Grover, N. Kambo, N. and Upadhyay, S.K., “ Kinetics and mechanism of palladium(II) catalysed oxidation of some alpha-aminoacids by chloramine-T in perchloric acid”, Ind. J. Chem., 41(A): 2482 (2012)
- [30] Lewis, E. S., Investigation of Rates and Mechanism of Reactions in Techniques of Chemistry, A Weissberger, 3<sup>rd</sup> ed., Wiley, New York, p. 421.(1974).
- [31] Onkar, A. S., Naik, P. N., Gunagi, S.D., Nandibewoor , S.T. and Chimatadar, A.S., “Kinetics and mechanism of oxidati of tyrosine by vanadium(V) in aqueous hydrochloric and ethanoic acid medium”, Ind. J. Chem., 51A, 1574(2012).
- [32] Gopalakrishanan, G. Pai, B.R. and Venkatasubramaniam, N., “Oxidation of alcohols by N-Bromosuccinimide-A curious dependence on initial concentration of the oxidant”, Ind. J. Chem, 19B, 293 (1998).
- [33] Chang, J.C. and Garner, G.S. “Kinetics of aquation of Aquopentachloroiridate(III) and Chloride anation of Diaquotetrachloroiridate(III) anions”, Inorg. Chem., 4(2), 209(1965).



10.22214/IJRASET



45.98



IMPACT FACTOR:  
7.129



IMPACT FACTOR:  
7.429



# INTERNATIONAL JOURNAL FOR RESEARCH

IN APPLIED SCIENCE & ENGINEERING TECHNOLOGY

Call : 08813907089  (24\*7 Support on Whatsapp)