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# A Chemical Synthesis of CdSnO<sub>3</sub> Nanoparticles and their Performance in Hydrogen Sensing

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**Abstract:** CdSnO<sub>3</sub> nanoparticles were successfully synthesized without any templates by simple co-precipitation synthesis route. Further characterized by using X-ray diffraction (XRD) measurements, field emission scanning electron microscopy (FESEM) and Fourier transform infrared (FTIR) spectroscopy and their hydrogen sensing properties were investigated. The CdSnO<sub>3</sub> nanoparticles exhibited outstanding gas sensing characteristics such as, higher gas response, extremely rapid response, fast recovery, excellent repeatability, good selectivity and at ambient operating temperature (~ 30°C). Furthermore, the CdSnO<sub>3</sub> nanoparticles can detect up to 5 ppm for hydrogen with reasonable sensitivity at an ambient operating temperature.

**Keywords:** Cadmium stannate, Hydrogen sensing, XRD, FTIR, TEM

## I. INTRODUCTION

In last two decades zero- and one-dimensional metal nanostructures, such as ZnO, TiO<sub>2</sub> and SnO<sub>2</sub>, have attracted enormous interest due to their unique properties and potential use in various applications such as photo catalysis, solar cells and gas sensors [1–5]. However, with keen research in nanotechnology, there is a demanding requirement for specially designed metal oxides to better match the properties of emerging materials. This has led to transformed interesting ternary metal oxides of the form ABO<sub>3</sub> such as cadmium stannate, zinc stannate and some metal titanate. Amongst these all metal, titanate is widely studied due to its piezoelectric nature while study of metal stannate less as compared to metal titanate. Because of this metal stannate having high electron mobility [6], high electrical conductivity, are chemically more stable than binary metal oxides and attractive optical properties that makes it suitable for a wide range of applications in solar cells, sensors for the detection of humidity and gases, negative electrode material for battery and as a photo catalyst[7-9]. The sizes and shapes of nano structures are crucial as it may affect their overall properties. Therefore, synthesis of nanostructures has considerably progressed over the last decade, to achieve different variety of shapes of nano materials. However, the synthesis of complex or ternary structures still remains a challenge for researchers. CdSnO<sub>3</sub> is *n*-type semiconductors with a band gap of 2–3 eV, which, due to a high concentration of native defects, are characterized by a rather high conductivity among all metal stannate [10]. The aim of present article is synthesizing CdSnO<sub>3</sub> nanoparticles by co-precipitation method. Further, characterize by various characterization techniques and evaluate the hydrogen sensing performance of CdSnO<sub>3</sub> thin film. As per our knowledge hitherto no body studied the hydrogen sensing properties of CdSnO<sub>3</sub> thin film.

## II. EXPERIMENTAL

### A. Materials

All chemicals were of analytical grade. The cadmium acetate, stannous chloride and sodium hydroxide were purchased from E-Merck (India) and were used without further purification.

### B. Synthesis of the CdSnO<sub>3</sub> nanoparticles

In this work, the CdSnO<sub>3</sub> nanoparticles were synthesized without any templates by using cadmium chloride, stannous chloride and ammonia as starting materials through a simple and low cost co-precipitation synthesis route. The cadmium chloride was used as the source of Cd<sup>2+</sup>, the stannous chloride was used as the source of Sn<sup>2+</sup> and ammonia was used as the precipitating agent to release hydroxyl ions slowly during the reaction. In a typical experiment, the aqueous solution containing 0.2 M Cadmium chloride, 0.6 M stannous chloride and 15 ml ammonia (30%) dissolved in 15 ml distilled water was prepared and added drop wise in the mixture of cadmium chloride and tin chloride to maintain the pH of the solution ~ 7 during the reaction and continuously stirred for 1 hour at room temperature 30 °C to obtain white coloured precipitate. The resulting cadmium hydroxyl stannate powder was washed with double distilled water and alcohol several times to remove impurities and by products present in the product.

The precipitate thus formed was dried at 80 °C in hot air oven for 12 h and grounded into a fine powder, which was then calcinated in air at different calcinating temperature at 400 °C for 2 h to obtain the end product.

C. Hydrogen Sensing Properties

The cadmium stannate nanoparticles powder was spin coated on the alumina substrate and the ohmic contacts were made with the help of silver paste to form gas sensing element. For the preparation of spin coated cadmium stannate thin films the cadmium stannate powder was dissolved in mixture of acetyl-acetone and ethanol in the ratio 8:2 to form a suspension in which 0.1 gm of p-hydroxy benzoic acid was added and the suspension was sonicated for one hour. The mixed suspension of cadmium stannate was spin coated using spin coater (SPN2000, Milman Thin Film Systems, Pvt. Ltd., Pune, India) forms the paste then the paste was coated on the alumina electrode and heated at 800°C to remove water from the film for the hydrogen sensing study. The electrical contact leads were fixed 0.7 cm apart with the help of silver paste on the surface of the film. The electrical resistance of the film was measured as a function of gas response by using a simple two probe configuration with a sensitive digital multimeter (2000 Digital multimeter, Keithley) controlled by a personal computer. The continuous variation in resistivity in the present of hydrogen gas was achieved in a simple experimental set-up fabricated in our laboratory in order to investigate the hydrogen sensing properties.

III. RESULTS AND DISCUSSION

A. Characterizations

The sensor was fabricated to study the different gas sensing properties of CdSnO<sub>3</sub>. The XRD pattern of as-prepared product annealed at temperature 400°C is depicted in Fig.1. All the diffraction peaks in the XRD pattern shown is indexed to cadmium stannate (JCPDS No.: 34-0885), indicating the formation of orthorhombic crystal structure (space group: Pb nm (62), a=5.4578, b=5.5773, c=7.8741) of distorted perovskite type structure. No other peaks were observed, indicating that no impurities were present and confirming that the adopted synthesis method gives pure CdSnO<sub>3</sub> nanoparticles. The average crystallite size was calculated by fitting the [2 0 0] diffraction peak (2θ = 32.9°) with a Gaussian function and using the values of the diffraction angle and peak full line width at half of maximum (FWHM) in the Debye-Scherrer formula –

$$D = \frac{k\lambda}{B \cos \theta} \dots\dots\dots(1)$$

where D is the average size of the crystallite, assuming that the grains are spherical, k is constant and it is ~ 0.9, λ is the wavelength of the X-ray radiation, B is the peak FWHM in radian and θ is the diffraction peak position. X-ray diffraction (XRD) analysis was performed with a Bruker diffractometer (D8, Advance, Bruker AXS model) with CuKα radiation (λ=1.5406 nm) operating at 40 kV and 40 mA. The average crystallite size of the CdSnO<sub>3</sub> nanoparticles was found to be in the range of 3.65 nm at 400°C. Wang et al and Meena et al shows the synthesis of CdSnO<sub>3</sub> nanoparticles and its crystal fitted by JCPDF data card, PDF#34-0758 which is rhombohedral, Hexagonal, R-3(148). They reported particle size of CdSnO<sub>3</sub> nanoparticles ~40-50 nm. Whereas Jia et al reported orthorhombic, β-CdSnO<sub>3</sub> nanoparticles has crystallite size 50 nm.

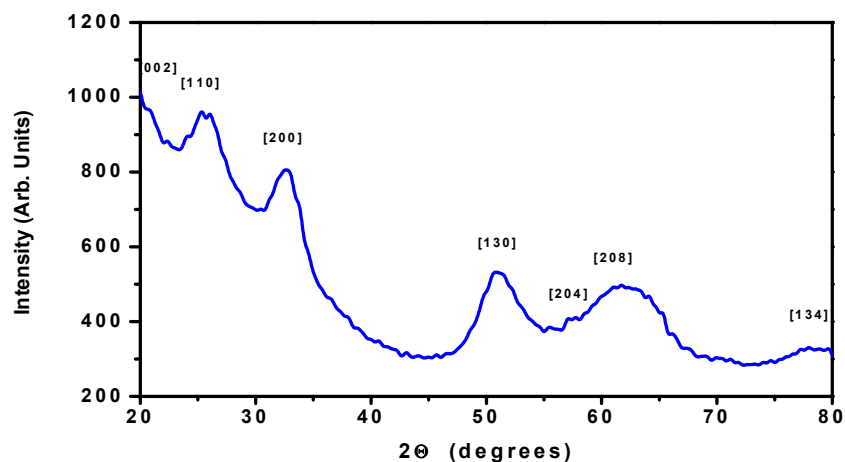


Fig. 1: XRD pattern of CdSnO<sub>3</sub> powder (calcinated ~400°C) prepared using Co-precipitation synthesis method.

The FTIR spectrum of the nanostructured CdSnO<sub>3</sub> heated at 400°C are shown in Fig.2. The FTIR spectrum for calcinated perovskite CdSnO<sub>3</sub> sample at 400°C is exhibits a broad band of which is mixing of three non-significant maxima of absorption between 630 and 690 cm<sup>-1</sup>, the first peak at 638 cm<sup>-1</sup> (Sn–O bond stretching along the *b* axis), the second at 661 cm<sup>-1</sup> (Sn–O bond stretching along the *a* axis, the 654 cm<sup>-1</sup> peak is very small), and the third at 687 cm<sup>-1</sup> (Sn–O bond stretching along the *a+c* direction). Mainly, the wide band at 440 cm<sup>-1</sup> is due to Sn–O–Sn scissoring. The peaks at 800–1400 cm<sup>-1</sup> are assigned to CdO. The bands in the region of 520–670 cm<sup>-1</sup> can be ascribed to the stretching vibration of Sn–O. An upward shift in the frequency range ~ 528–551 cm<sup>-1</sup> is due to the presence of Sn in Cd–O lattice [11-12, 19]. The FTIR spectroscopy analysis was performed with a Nicolet FTIR spectrometer (IMPACT 420 DSP) by the conventional KBr method in the spectral range 4000–400 cm<sup>-1</sup>

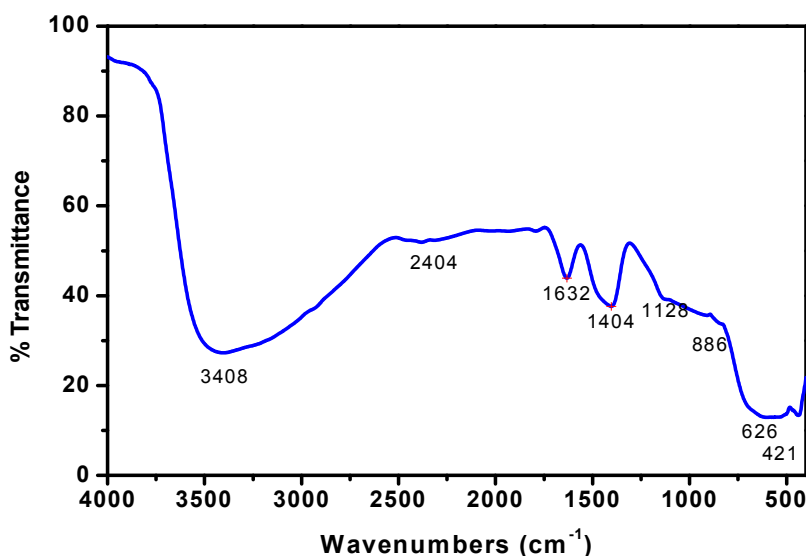


Fig. 2: FTIR spectrum of CdSnO<sub>3</sub> nanoparticles calcinated at 400°C

The TEM image of as-prepared product [Fig.3(a)] exhibits a non-uniform shaped, narrow sized distributed and agglomerated the nanoparticles at 400°C. The average grain size of the CdSnO<sub>3</sub> nanoparticles is estimated to be around 5-6 nm, which is nearly matches to XRD crystallite size. Fig. 3 (b) shows the high resolution TEM (HRTEM) image. This HRTEM image shows non-uniform fringes and drastic variation in intensity region-wise. This study suggests that nanoparticles are non-uniformed sizes and oriented in particular direction like a single crystalline structure. The surface morphological study was performed by a high-resolution transmission electron microscope (HRTEM, Tecnai G2 20 Twin, FEI, USA)

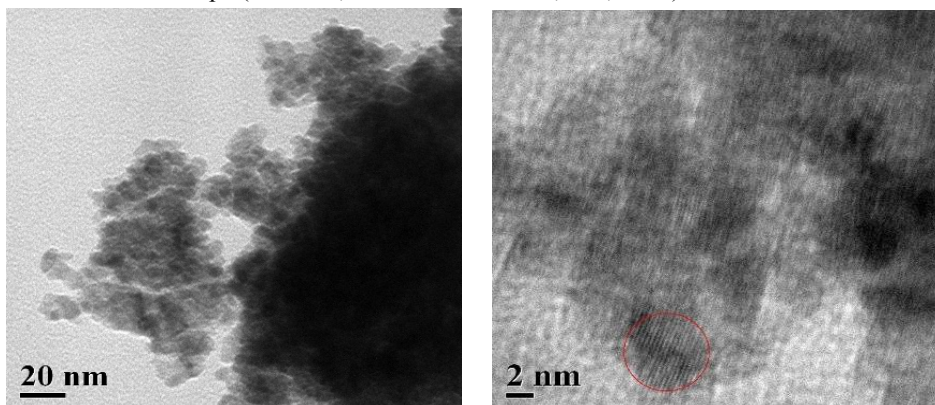


Fig.3: (a) TEM (b) HRTEM image of 400°C calcinated CdSnO<sub>3</sub> nanoparticles

**B. Hydrogen Sensing Characteristics**

The H<sub>2</sub> gas sensing experiments were performed at different temperatures in order to find out the optimum operating temperature for H<sub>2</sub> gas detection. Before exposing to the H<sub>2</sub> gas, the sensing element could equilibrate inside the gas chamber at an operating temperature for 1 h. The effect of an operating temperature on the gas response of CdSnO<sub>2</sub>nanoparticle based sensor to 50 ppm H<sub>2</sub> is shown in Fig.4.

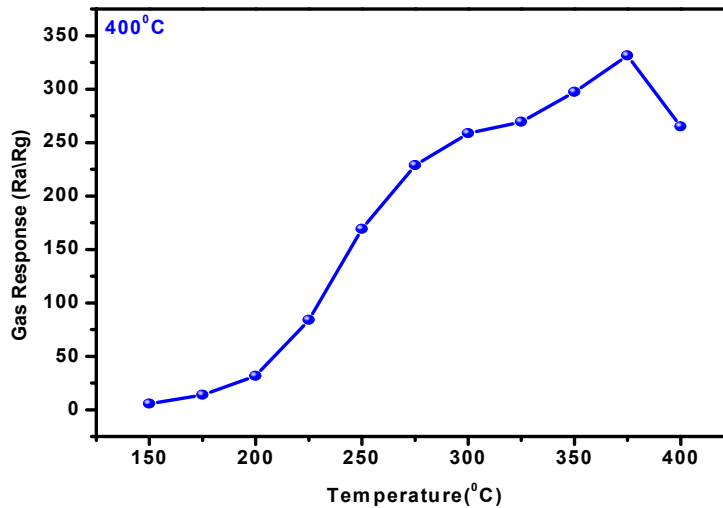


Fig.4: Effect of operating temperature on the gas response of CdSnO<sub>3</sub> nanoparticles (calcinatad at 400°C) powder to 50 ppm H<sub>2</sub> gas.

The relationship between the gas response and the operating temperature exhibits a trend of “increase-maximum-decay” behaviour to 50 ppm H<sub>2</sub> gas. To investigate the various H<sub>2</sub> sensing characteristics of this sample such as response and recovery, reproducibility and selectivity, operating temperature is optimized to 300°C. The response and recovery characteristics of the CdSnO<sub>3</sub> nanoparticles to 50 ppm H<sub>2</sub> gas at an operating temperature 375°C is shown in Fig. 5. It was observed that the resistance of the sensing element decreases when exposed to the H<sub>2</sub>. As can be seen from Fig. 5, the sensor responds very rapidly after introduction of H<sub>2</sub> and recovers slowly when it is exposed to air. The CdSnO<sub>3</sub> nanoparticles have response time of ~ 2-3 s and the recovery time of ~ 15-17 s. The CdSnO<sub>3</sub> nanoparticles show good reproducibility and reversibility upon repeated exposure and removal of H<sub>2</sub> under same conditions. This suggests that the CdSnO<sub>3</sub> nanoparticles can be used as a reusable sensing material for the detection of H<sub>2</sub>.

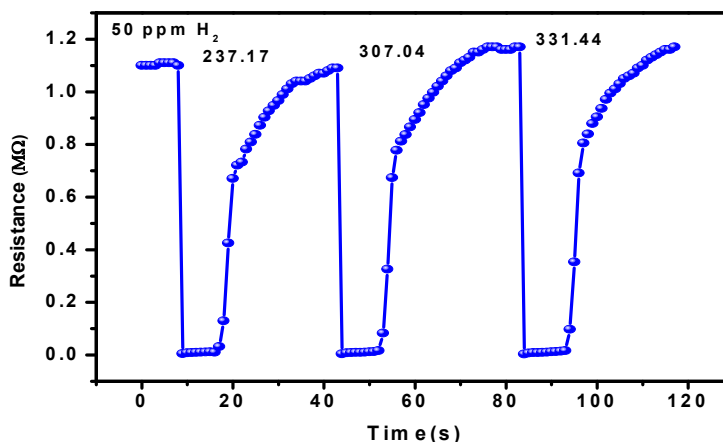


Fig.5: Repetitive response of CdSnO<sub>3</sub> nanoparticles to 50 ppm H<sub>2</sub> gas at an operating temperature of 400°C.

The gas response of the CdSnO<sub>3</sub> nanoparticles versus H<sub>2</sub> gas concentration at an operating temperature of 375°C is shown in Fig. 6. It was observed that the gas response increases linearly in the range 5-50 ppm H<sub>2</sub> gas. It is found that the response of CdSnO<sub>3</sub> nanoparticles can be empirically represented as,

$$y = -12.72 + 7.072 * x, R_2 = 0.99224,$$

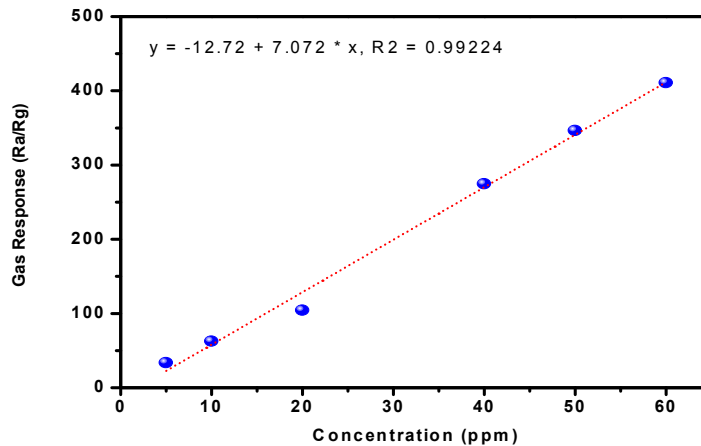


Fig.6: Response of CdSnO<sub>3</sub> nanoparticles to 50 ppm H<sub>2</sub> gas at an operating temperature of 400°C.

Where x, y and R<sub>2</sub> represents the H<sub>2</sub> concentration, gas response and correlation coefficient, respectively. The dotted line shows the linear fit to the experimental data, illustrating clearly good quality of the fit. The linear relationship between the gas response and the H<sub>2</sub> concentration at low concentrations (5-50 ppm) may be attributed to the availability of enough sensing sites to act upon the CdSnO<sub>3</sub> nanoparticles.

Selectivity is an important parameter of gas sensors and it is the ability of a sensor to respond to a certain gas in presence of other gases. Theoretically, the sensors should have high response to some gases and little or no response to other gases in the same surroundings. To study the selective behaviour of the CdSnO<sub>3</sub> nanoparticles to H<sub>2</sub> at an operating temperature of 300 °C, the gas response towards LPG, CO, CO<sub>2</sub> and ethanol with concentration 50 ppm each were also measured.

The selectivity property of CdSnO<sub>3</sub> nanostructured thin film at various pollutant gases is shown in Fig. 7. The CdSnO<sub>3</sub> nanoparticles exhibit higher response to H<sub>2</sub> (331), whereas it shows a considerably lower response (<7.62) to LPG, CO, CO<sub>2</sub> and ethanol. The selectivity coefficient (K) of H<sub>2</sub> to another gas is defined as [13, 14]:

$$K = \frac{S_{H_2}}{S_B}$$

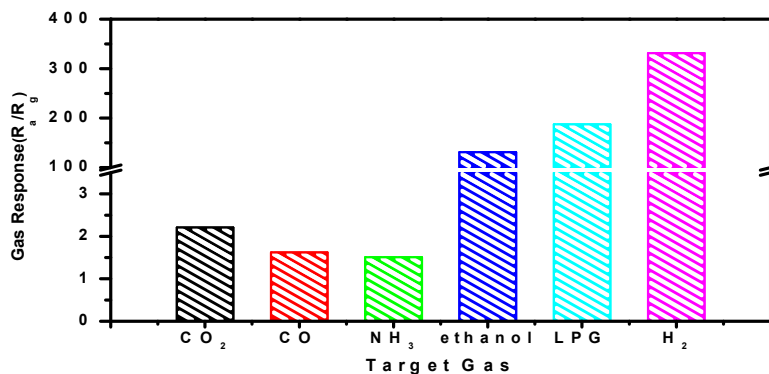


Fig. 7: Bar chart showing the gas response of CdSnO<sub>3</sub> nanoparticles for different gases. The gas concentration and operating temperature in all cases were 50 ppm and 400°C, respectively.

Here  $S_{H_2}$  and  $S_B$  are the responses of sensors in  $H_2$  and B gases, respectively. The selectivity coefficients for the  $CdSnO_3$  nanoparticles were 187.21 to LPG, 2.21 to  $CO_2$ , 1.62 to CO and 131.21 to ethanol. The experimental results indicate that the  $CdSnO_3$  nanoparticles-based sensor has a good selectivity to  $H_2$ . The reproducibility and stability of the  $CdSnO_3$  nanoparticles were measured by repeating the measurement many times.

In literature possible hydrogen sensing mechanism are explained on the basis of adsorption and desorption mechanisms. The probable gas sensing mechanism is upon exposure to  $H_2$  gas, much greater number of trapped electrons are released once the adsorbed surface O species are chemically reduced by the  $H_2$  molecules leading to lowering of the barrier height and increasing the conductivity. Therefore, change in the current as well as the magnitude of the highest current in  $CdSnO_3$  nanoparticles is much higher resulting in the enhanced gas response. The Cd involves in Sn-O lattice is already observed in FTIR and XRD result. An increase in the surface-to-volume ratio of the  $CdSnO_3$  nanoparticles which would increase number of adsorbed O molecules and an increase in the surface defects, which can influence the chemical as well as electronic properties, the adsorption behaviour. Further these defects also control the carrier concentration via effective near surface electron depletion. Thus, these are main reasons enhancement of  $H_2$  gas response [15-18].

#### IV. CONCLUSIONS

- 1) We have successfully synthesized the  $CdSnO_3$  nanoparticles at low cost by using a simple co-precipitation method by calcination of  $CdSnO_3$  nanoparticles.
- 2) XRD and FTIR results clearly indicating formation of  $CdSnO_3$  nanostructure. TEM study revealed the formation of single crystalline nanostructures.
- 3) The gas response to 50 ppm of  $H_2$  gas is found to be  $\sim 331.44$ . The response time was nearly 3-4 sec and the recovery time was found to be 5-6 sec.
- 4) The synthesized  $CdSnO_3$  nanoparticles are able to detect up to 5 ppm for  $H_2$  with reasonable response at room temperature. Further, it was shown that the  $CdSnO_3$  nanoparticles can be reliably used to monitor the concentration of  $H_2$  gas over the range (5-60 ppm).
- 5) These results indicate that the  $CdSnO_3$  nanoparticles are indeed very attractive  $H_2$  gas sensing materials.

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